T09D03 – (9.5) Electrolytic Cells Notes

Name…….

1. 9.5.1 Describe, using a diagram, the essential components of an electrolytic cell. (2)
   a. Diagram a simple Electrolytic Cell

2. 9.5.2 State that oxidation occurs at the positive electrode (anode) and reduction occurs at the negative electrode (cathode). (1)
   a. Define an electrolyte

   b. During electrolysis, we still follow RedCat and AnOx, what does this mean?

3. 9.5.3 Describe how current is conducted in an electrolytic cell. (2)
   a. How do we test conductivity?

   b. What does it mean for a substance to be a conductor or an insulator?

4. 9.5.4 Deduce the products of the electrolysis of a molten salt. (3)
   a. What happens when an electrical current passes through an ionic substance (molten or solution)?

Diagram, and provide Redox half equations as well as the full equation, for the decomposition of $PbBr_2$	Diagram, and provide Redox half equations as well as the full equation, for the decomposition of NaCl
Diagram, and provide Redox half equations as well as the full equation, for the decomposition of H_2O	Diagram, and provide Redox half equations as well as the full equation, for electroplating a material with Cu